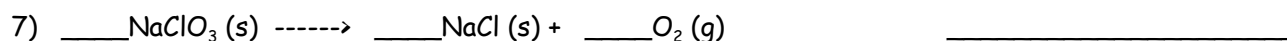
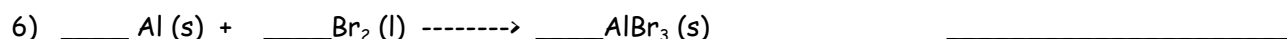
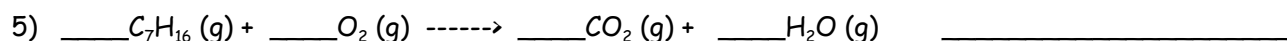
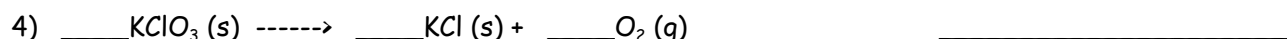
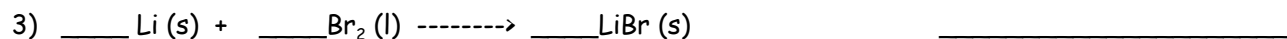
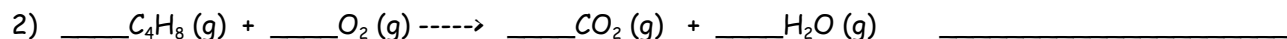


Equations #2

Name: _____

Part 1. Supply the correct coefficients to balance the following reaction equations (assume that all formulas are correct). Then classify each reaction as either synthesis, decomposition, or combustion.

Type of Reaction



Part 2. Write the correct formulas for all reactants and products, then supply the necessary coefficients to balance the equations. Then classify each equation according to its reaction type.

8) Magnesium metal reacts with liquid bromine to form liquid magnesium bromide.

9) Liquid pentane (C_5H_{12}) reacts with oxygen gas to form carbon dioxide gas and liquid water.

10) Solid barium carbonate can be broken down into solid barium oxide and carbon dioxide gas.

11) Hydrogen gas reacts with chlorine gas to produce hydrogen chloride (HCl) gas.

12) Liquid cyclohexene (C_6H_{12}) will burn in the presence of oxygen gas to form carbon dioxide gas and water vapor.

13) Solid potassium chlorate can be decomposed by heating into solid potassium chloride and oxygen gas.

Part 3: Identify each of the following reactions by writing the name of the reaction on the line to the left of the chemical reaction. Complete the reaction on the line to the right. Be sure to balance the equation.

Reaction Type	Reaction		
14) _____	_____ C_8H_8 + _____ O_2	-->	_____
15) _____	_____ Mg + _____ O_2	-->	_____
16) _____	_____ $CrCl_3$	-->	_____
17) _____	_____ Na + _____ Cl_2	-->	_____
18) _____	_____ C_3H_8 + _____ O_2	-->	_____
19) _____	_____ Ba + _____ N_2	-->	_____
20) _____	_____ Rb_2S	-->	_____