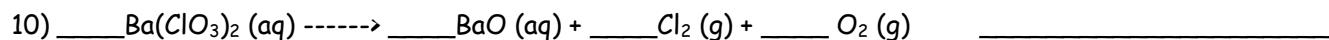
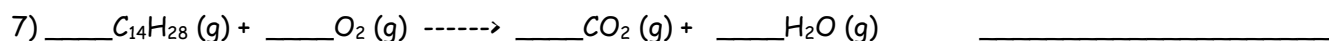
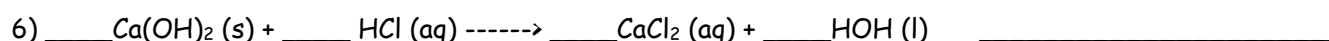
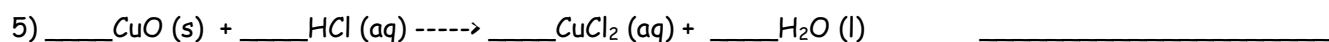
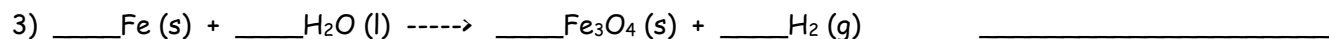
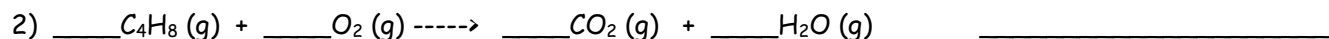
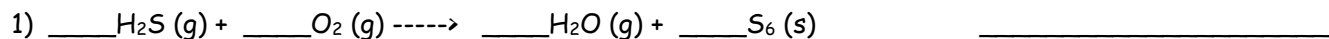


Equations #3

Name: _____

Part 1. Supply the correct coefficients to balance the following reaction equations (assume that all formulas are correct). Then classify each reaction as either synthesis, decomposition, single replacement, double replacement or combustion.



Part 2. Write the correct formulas for all reactants and products, then supply the necessary coefficients to balance the equations. Then classify each equation according to its reaction type.

11) Solid chromium metal reacts with liquid iodine to form liquid chromium (III) iodide.

12) Rubidium chloride and magnesium sulfate combine to produce rubidium sulfate and magnesium chloride, all in aqueous solution.

13) Aqueous aluminum sulfide reacts with hydrogen gas to produce aqueous hydrosulfuric acid (H_2S) and aluminum metal.

14) Using electrolysis, the compound dinitrogen pentoxide (N_2O_5) can be broken down into nitrogen gas and oxygen gas.