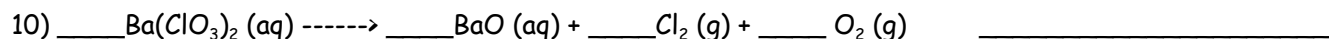
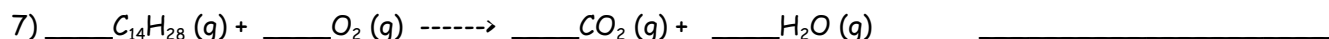
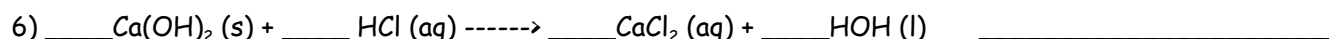
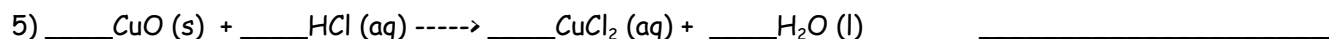
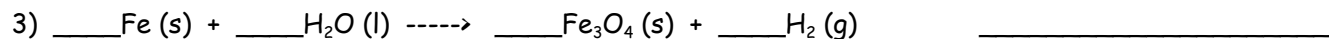
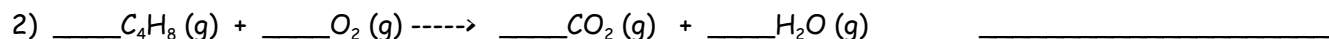
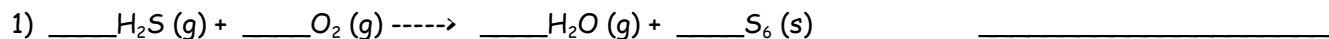


Equations #3

Name: _____

Part 1. Supply the correct coefficients to balance the following reaction equations (assume that all formulas are correct). Then classify each reaction as either synthesis, decomposition, single replacement, double replacement or combustion.



Part 2. Write the correct formulas for all reactants and products, then supply the necessary coefficients to balance the equations. Then classify each equation according to its reaction type.

11) Solid chromium metal reacts with liquid iodine to form liquid chromium (III) iodide.

12) Rubidium chloride and magnesium sulfate combine to produce rubidium sulfate and magnesium chloride, all in aqueous solution.

13) Aqueous aluminum sulfide reacts with hydrogen gas to produce aqueous hydrosulfuric acid (H_2S) and aluminum metal.

14) Using electrolysis, the compound dinitrogen pentoxide (N_2O_5) can be broken down into nitrogen gas and oxygen gas.

15) Liquid octane (C_8H_{18}) in gasoline burns in the presence of oxygen to form carbon dioxide gas and liquid water.

Part 3: Identify each of the following reactions by writing the name of the reaction on the line to the left of the chemical reaction. Complete the reaction on the line to the right. Be sure to balance the equation.

Reaction Type	Reaction	
16) _____	_____ CaBr_2 + _____ O_2	--> _____
17) _____	_____ NaCl + _____ AgNO_3	--> _____
18) _____	_____ Mg + _____ NiCl_3	--> _____
19) _____	_____ $\text{Pb}(\text{NO}_3)_2$ + _____ KBr	--> _____
20) _____	_____ AuF_3 + _____ Na	--> _____
21) _____	_____ Mg + _____ O_2	--> _____
22) _____	_____ $\text{C}_{11}\text{H}_{24}$ + _____ O_2	--> _____